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Intro

How many protons

Naming rules

Percent composition

Nitrogen gas

Oxidation State

Stp

Example

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Introduction

Volume Mass Percent

Mole Fraction

Molarity

Harder Problems

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convert the moles of substance a to the moles of substance b

convert it to the moles of sulfur trioxide

react completely with four point seven moles of sulfur dioxide

put the two moles of so2 on the bottom

given the moles of propane

convert it to the grams of substance

convert from moles of co2 to grams

react completely with five moles of o2

convert the grams of propane to the moles of propane

use the molar ratio

start with 38 grams of h2o

converted in moles of water to moles of co2

using the molar mass of substance b

convert that to the grams of aluminum chloride

add the atomic mass of one aluminum atom

change it to the moles of aluminum

change it to the grams of chlorine

find the molar mass

perform grams to gram conversion

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