Kineticheskaia Teoriia Protsessov Perenosa I Relaksatsii V Potokakh Neravnovesnykh Reagiruiushchikh Gazov

#kinetic theory #nonequilibrium gases #transport processes #relaxation processes #reacting gas flows

Explore the fundamental principles of kinetic theory as applied to transport and relaxation processes within flows of nonequilibrium reacting gases. This area delves into the complex molecular interactions, energy transfer, and chemical reactions that occur when gases are far from thermal equilibrium, offering crucial insights for fields such as combustion, plasma physics, and high-speed aerodynamics where accurate modeling of gas behavior under dynamic conditions is essential.

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Equation of State for an Ideal Gas and Ideal Gas Law
Important Formulae
Average Translational Kinetic Energy
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The Average Force
Second Law of Motion
Average Force Acting on the Wall
Find the Pressure
The Ideal Gas Law
Ideal Gas Law
Solve for the Average Translational Kinetic Energy
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Kinetic Energy in a Molecule
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Gas Constant
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Kinetic Energy of a Mole of Nitrogen Gas
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THE INTER MOLLECULAR FORCES OF ATTRACTION BETWEEN THE GAS MOLECULES ARE NEGLIGIBLE
THE MOLECULES OF THE GAS BEHAVE LIKE PERFECTLY ELASTIC SPHERES
THE VOLUME OCCUPIED BY MOLECULES OF THE GAS IS NEGLIGIBLE COMPARED TO THAT OF THE CONTAINER 3.e MOLECULES OF THE GAS ARE CONSIDERED TO BE POINT MOLECULES
THE DURATION OF COLLISION BETWEEN MOLECULES OF A GAS AND THE WALLS OF THE CONTAINER IS NEGLIGIBLE COMPARED TO THE TIME BETWEN COLLISIONS
THE MOLECULES MOVE ABOUT IN CONTINUED RANDOM MOTION
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Formula
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Dilute gas
Table
Discussion
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