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Which of the following will give a straight line plot in the graph of $\ln[A]$ versus time?

Which of the following units of the rate constant K correspond to a first order reaction?

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The initial concentration of a reactant is 0.738 M for a zero order reaction. The rate constant k is 0.0352 M/min . Calculate the time it takes for the final concentration of the reactant to decrease to 0.255 M .

Calculate the rate constant K for a second order reaction if the half life is 243 seconds. The initial concentration of the reactant is 0.325 M .

Which of the following particles is equivalent to an electron?

Identify the missing element.

The half-life of Cs-137 is 30.0 years. Calculate the rate constant K for the first order decomposition of isotope Cs-137.

The half life of Iodine-131 is about 8.03 days. How long will it take for a 200.0g sample to decay to 25g?

Which of the following shows the correct equilibrium expression for the reaction shown below?

Calculate K_p for the following reaction at 298K. $K_c = 2.41 \times 10^{-2}$.

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a net charge of positive to the right side

start with the first one

add 3 electrons to the side with a higher charge

add the two half reactions we need

add these two half-reactions

add six H^+ ions to the left

add 6 electrons to the left side

need to cancel the 6 electrons on both sides

check the total charge the

start by balancing it under acidic conditions

add four hydroxide ions to the left side

add the 3 electrons to the left side

add 4 water molecules on the right side

add eight hydroxide ions to both sides

produces 1 chloride ion and 8 hydroxide

the charges

add 8 electrons to the left

produce three chloride ions and 24 hydroxide ions

subtract both sides by 24 hydroxide ions

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Unboxing

chloride test

sulfate test

total alkalinity test

total hardness test

calcium hardness test

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Naming cations

Stock system

Question of the day

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Oil Drop Experiment

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Making Ice Cream

Cooking vs Chemistry

Atomic Mass

Atomic Mass Units

Isotopes

Quiz

Review

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Intro

How many protons

Naming rules

Percent composition

Nitrogen gas

Oxidation State

Stp

Example

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Intro

Split Up Into Half-Reactions

Balance H with Hydrogen Ions

Balance Charge With Electrons

Combine the Half-Reactions

Balance Elements other than H and o

Balance O With Water Molecules

Make The Electron Numbers Equal

Add Hydroxides to Both Sides

Combine Protons And Hydroxides

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Final Thoughts

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A Horse
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