

## chemistry notes for class 11 maharashtra board

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Question 4.1

Question 4.2

Question 4.3

Question 4.4

Question 4.5

Question 4.6

Question 4.7

Question 4.8

Question 4.9

Question 4.10

Question 4.11

Question 4.12

Question 4.13

Question 4.14

Question 4.15

Question 4.16

Question 4.17

Question 4.18

Question 4.19

Question 4.20

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Question 4.35

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Question 4.37

Question 4.38

Question 4.39

Question 4.40

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Introduction

NCERT Q 4.1

NCERT Q 4.2

NCERT Q 4.3

NCERT Q 4.4

NCERT Q 4.5

NCERT Q 4.6

NCERT Q 4.7

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Introduction

Exercise 4.1

Exercise 4.2

Exercise 4.3

Exercise 4.4

Exercise 4.5

Exercise 4.6

Exercise 4.7

Exercise 4.8

Exercise 4.9

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Exercise 4.34

Exercise 4.35

Exercise 4.36  
Exercise 4.37  
Exercise 4.38  
Exercise 4.39  
Exercise 4.40

Class 11th Chemistry Chapter 7 | Exercise Questions (7.1 to 7.30) | Redox Reactions | NCERT - Class 11th Chemistry Chapter 7 | Exercise Questions (7.1 to 7.30) | Redox Reactions | NCERT by Ignited Minds 91,920 views 3 months ago 3 hours, 4 minutes - This video includes a detailed explanation of exercise questions of **chapter**, 7 (Redox Reactions). If you want to view a particular ...

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Question 7.2  
Question 7.3  
Question 7.4  
Question 7.5  
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Question 7.30

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Introduction

Atoms

Molecules formation theories

Kossel-Lewis approach

Lewis Symbols

Octet Rule

Lewis Dot Structures

Lewis Dot Structure: CO<sub>2</sub>

Lewis Dot Structure: H<sub>2</sub>O

Lewis Dot Structure: C<sub>2</sub>H<sub>4</sub>

Lewis Dot Structure: C<sub>2</sub>H<sub>2</sub>

Lewis Dot Structure: Carbonate ion

## Why F.C

Limitation: Octet Rule  
Ionic Bond  
I.E.  
Electron Gain Enthalpy  
Lattice Enthalpy  
Ex: NaCl  
Stable Ionic Bond  
Bond Length  
Atomic radius  
Vander Waals Radii  
Bond Enthalpy  
Bond Angle  
Bond Order  
Resonance Struc.  
Bond Polarity  
D.M: Polyatomic  
Fajan's Rule  
VSEPR Theory  
Postulate 1  
Postulate 2  
Postulate 3  
Postulate 4  
Geometry VSEPR  
Shape vs. Geometry  
VSEPR theory: CO<sub>2</sub>  
VSEPR: Merits  
VSEPR: Limitations  
VBT  
Orbital Overlap  
Orbital Overlap Types  
Covalent Bond types  
Sigma Bond  
Pi bonds  
Str.: Å & Å  
Rotation  
Å/s. Å  
Draw Shapes  
Hybridisation  
Hybridisation: sp  
Hybridisation: sp<sup>2</sup>  
Hybridisation: sp<sup>3</sup>  
Hybridisation: d orb.  
VBT: Limitations  
MOT  
MOT formation  
LCAO conditions  
MO: Types  
Energy Dia. : MO  
Molecular Behaviour  
Bond Order  
Stability Molecules  
Bond Length  
Magnetic nature  
Hydrogen Bonding  
Covalent Vs. H bond

Chemical Bonding | NCERT Solutions for Class 11 Chemistry - Dr. Prashant Sir | Vedantu JEE -  
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Srivastava Sir will be covering all the **NCERT Solutions for Class 11 Chemistry Chapter 4**, -



Chemical ...

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Introduction :Chemical Bonding & Molecular Structure

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Ques 18 to 23

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intro

Octet Rule

Types of Bond

Octet Rule Limitation

Bond Pair & Lone Pair

Method 1 ( Find Structures )

Method 2 ( Formula Method )

Most Stable Structure

Valence Bond Theory (VBT)

Limitation in VBT

Molecular Orbital Theory

Paramagnetic vs Diamagnetic

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